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# Handheld femtogram detection of hazardous Picric acid with hydrophobic Ag nanopillar SERS substrates and mechanism of elasto-capillarity

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*Supporting Information Available: The following files are available free of charge: Movie S1, MD simulation of the wetting; Movie S2, MD simulation of the evaporation; Table S1, Raman peaks*

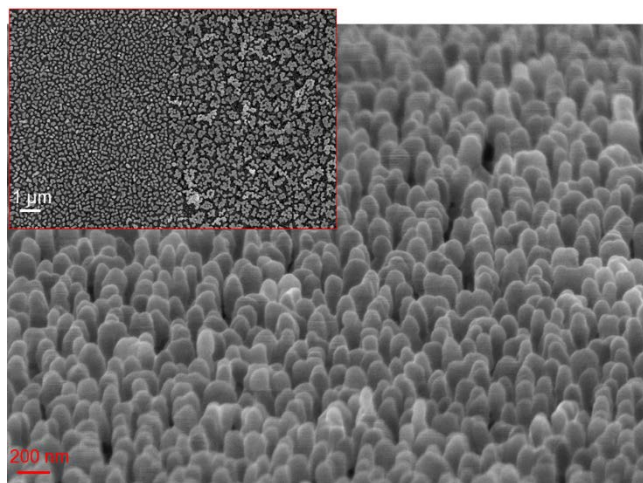
**ABSTRACT:** Picric acid (PA) is a severe environmental and security risk due to its unstable, toxic and explosive properties. It is also challenging to detect in trace amounts and in-situ because of its highly acidic and anionic character. Here we assess sensing of PA under non-laboratory conditions using surface-enhanced Raman scattering (SERS) silver nanopillar substrates and hand-held Raman spectroscopy equipment. The advancing elasto-capillarity effects are explained by molecular dynamics simulations. We obtain a SERS PA detection limit on the order of 20 ppt, corresponding to attomole amounts, which together with the simple analysis methodology demonstrates that the presented approach is highly competitive for ultrasensitive analysis in the field.

Picric acid (PA) or 2,4,6-trinitrophenol (TNP), is an old and well-known high explosive but mounting security concerns has led to a renewed interest and an outburst of publications, for example.<sup>1</sup> PA is highly unstable, in addition to being toxic and cancerous and, therefore, a serious security liability.<sup>1a-v</sup> There is an urgent need for highly sensitive PA sensors for water samples since large amounts have been released from industrial use, leading to accumulation in rivers, lakes and ground water.<sup>1a-v</sup> The marketing and use of some explosive precursors has recently been regulated within

the European Union (EU No 98/2013) to prevent illicit manufacture. However, some explosives, including PA, can still be synthesized from freely available chemicals. PA is a solid with low vapour pressure and a relatively high water solubility (12.7 g/l), which makes it difficult to detect in vapour phase at ambient temperatures. PA present as dust is classified as a sensitizing substance and has recently been assigned a new low permissible exposure limit (0.1 mg/m<sup>3</sup> 8 hr TWA). To this end, robust, sensitive and selective PA detection methods need to be developed, both to avert terrorist threats as well as for monitoring environmental pollution.<sup>1</sup>

Most recent works on PA have dealt with the issue of trace detection and sensing, in particular based on fluorescence techniques.<sup>1</sup> Surface-enhanced Raman scattering (SERS) spectroscopy is widely recognized as an ultrasensitive method to detect e.g. biomolecules<sup>2</sup> and explosives,<sup>3</sup> and in particular conjugated nitro-explosives such as TNT and PA.<sup>3b</sup> Despite this, the first SERS studies of PA appeared only recently.<sup>3c, 4 5</sup> For example, Stewart et al.<sup>5</sup> demonstrated a detection limit of 20 µM PA on positively charged Ag nanoparticles and pointed out the difficulty of measuring anions, such as PA, on SERS active gold or silver due to the negative zeta potential of these metals. The lowest SERS detection limit for PA, 0.46 µM using a hybrid silver- magnetite detection scheme, was reported recently.<sup>4</sup>

We have recently reported that gold-on-silicon nanopillar SERS substrates exhibit extraordinary properties for detecting the nerve gases VX and Tabun.<sup>6</sup> We here use similar substrates in which the Au has been replaced by silver. This type

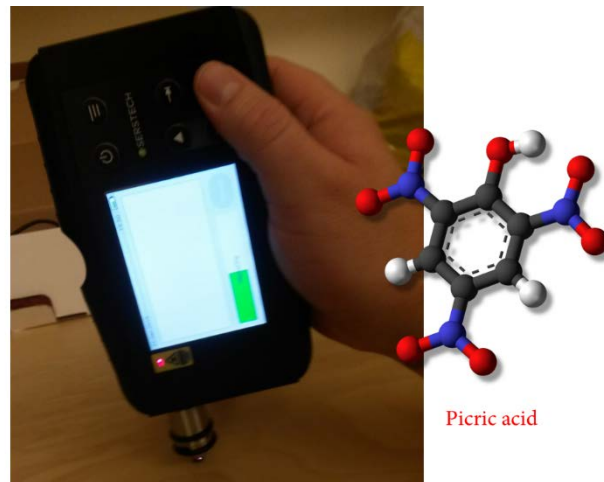


**Figure 1.** Tilted SEM image of the silver nanopillar SERS substrates used in this study. Inset shows the substrate at an evaporated droplet interface, where the droplet has generated clustering to the right.

of substrate has been described previously<sup>7</sup> and further optimized.<sup>8</sup> Briefly, the samples consist of a Si-nanopillar base substrate, obtained through reactive ion etching, covered by 225 nm evaporated Ag. The resulting “nano-forest”<sup>9</sup> of flexible Ag nanopillars, shown in Figure 1, turned out to be less hydrophobic than the gold substrates. Hence, it was easier to place and hold small water based droplets on them with good precision. The droplets almost immediately adhered, similar to the Au-nanopillar case, with high contact angles ( $120 - 140^\circ$ ) characteristic of hydrophobic surfaces. If a droplet is placed on a substrate and left to evaporate, elasto-capillary forces pulls the nanopillars together, resulting in SERS hot-spot clusters.<sup>6</sup> This is illustrated by the SEM image inset in Figure 1, which shows the edge between a pristine surface region and a region where a droplet containing a low concentration of PA (0.4 ppb) has dried out. Similar hot-spot generating clustering has been observed previously on several gold and silver substrates, e.g. in.<sup>10</sup> The clustering effect also creates a macroscopically visible spot on the substrates (ca. 1.5 mm diameter from a 2  $\mu$ l droplet). Such spots can be clearly seen on silver-based substrates under normal lighting conditions, whereas Au-substrates require very bright light. Figure 2 illustrates how we used a handheld Raman instrument (Serstech 100 Indicator) to measure SERS from PA (molecular model in Fig. 2 inset) on the Ag nanopillar substrates. Rapid droplet adhesion, clearly visible drying stains and the possibility to perform handheld measurements is clearly a powerful and convenient combination for SERS analysis in the field and outside dedicated laboratories.

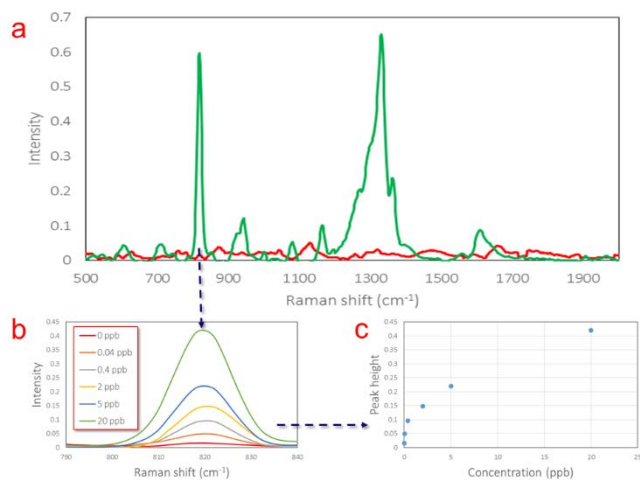
The SERS spectrum in figure 3a from an evaporated 2  $\mu$ l droplet with 20 ppb PA in water show several distinctive peaks, of which the major ones are at: 820 and 1332  $\text{cm}^{-1}$  and are both related to vibrational modes of the  $\text{NO}_2$ -group.<sup>11</sup> In table S1 all identified peaks are listed and compared with literature SERS and Raman studies,<sup>5, 11a</sup> and a high correlation is seen, especially with the SERS study. Fig. 3b shows the 820

$\text{cm}^{-1}$  peak at different concentrations of Picric acid (0 – 20 ppb, averaged spectra,  $n_{\text{Tot}} = 22$ ). The measurements were acquired as shown in Fig. 2, with experimental parameters: 785 nm/ 165 mW laser, 2 s acquisition time, 50  $\mu\text{m}^2$  spotsize, numerical aperture of the outermost lens was 0.3. A nice trend is seen in the calibration curve (Fig. 3c), and as typically seen in SERS a rapidly declining signal, which as commonly, is likely due to hot-spot saturation.<sup>12</sup> The lowest explicitly measured concentration was 0.04 ppb, and the 2  $\mu$ l droplet equals to a total amount

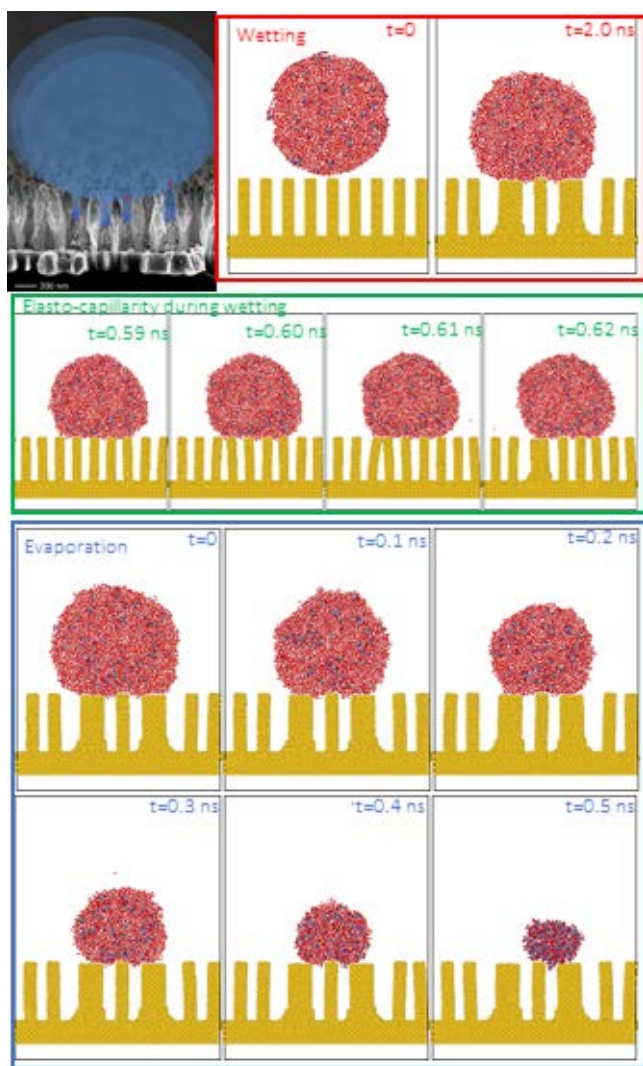


**Figure 2.** The Raman instrument during handheld measurement on a SERS substrate and a molecular model of Picric acid.

of 80 femtogram or 350 attomol. From zero level noise ( $\sigma$ ) and the slope ( $k$ ) to the 0.04 ppb level the detection limit ( $\text{LOD} = 3\sigma/k$ )<sup>13</sup> were calculated to be 18 ppt (parts per trillion; total amount 36 femtogram, 160 attomol), and limit of quantification (LOQ) to 0.06 ppb. For a brief comparison with literature, the previously mentioned SERS study showed a LOD of 0.46  $\mu\text{M}$  which equals about 110 ppb, i.e. approximately 3 – 4 orders of magnitude higher. Looking at other techniques a very impressive LOD was demonstrated recently by Malik et al.,<sup>1m</sup> 7 ppt or 31 pM and the lowest explicitly detected PA concentration was 200 pM or 46 ppt, i.e. very similar results as in present study (18 ppt and 40 ppt). The technique used by Malik et al. was fluorescence quenching based, which is by far the most used sensing method for Picric acid. Additionally, for our most impressive results regarding totally loaded amounts of PA comparing numbers are not present in the Malik et al. study. For a brief comparison with the in-laboratory gold standard for everything difficult to measure mass-spectrometry, which quite commonly achieves detection limits in the femtogram range, however, with a quick overview of literature picogram detection of Picric acid was the best to be found.<sup>14</sup>



**Figure 3.** a) Handheld instrument Serstech SERS spectrum of Picric acid (2 µl, 20 ppb) and background (red). b) Concentration dependence of the 820 cm⁻¹ peak. c) PA calibration curve.



**Figure 4.** Top: profile SEM image of the Ag nanopillar substrates. Illustrated on top is an evaporating water droplet. Red panel: MD snapshots of the droplet of PA solution wetting on Ag pillared surface. Balls with different colors were used to denote different type of atoms: Oxygen (red);

Hydrogen (white); Carbon (grey); Nitrogen (blue); Ag (yellow). Left: Initial state; Right: Equilibrium state at 2.0 ns. Green panel: clustering of two Ag nanopillars due to the elasto-capillarity interactions. Blue panel: MD snapshots of evaporation of a droplet of PA solution on Ag nanopillars.

Further insights into the nanofluidics of PA/water droplets on Ag nanopillar surfaces were investigated by molecular dynamics (MD) simulations.<sup>15</sup> MD simulations were carried out to study the wetting and evaporation process on an Ag nanopillared surface, as shown in Fig. 4. Each Ag nanopillar has a height of 4.09 nm and a width of 0.82 nm and the Ag-Ag interactions are modelled by (EAM) potential.<sup>16</sup> 5,000 water molecules and 100 PA molecules were mixed together using Packmol.<sup>17</sup> Water TIP4P/2005 model was used<sup>18</sup> and PA was described by the Merck Molecular ForceField.<sup>19</sup> The temperature of the system was maintained at 300 K using a Berendsen thermostat. A timestep of 1.0 fs was used. All the MD simulations were performed using LAMMPS.<sup>20</sup> We also calculated the surface tension of water and PA solution using Kirkwood-Buff formula.<sup>21</sup> Our MD results show that the surface tension for pure water and PA/water solution are 62.8 and 72.5 mN/m, respectively, i.e. the PA molecules increases the surface tension compared with pure water.

The droplet of PA solution, with a diameter of about 9.0 nm, was brought to connect with the pillared surface and then started to wet the surface. Due to the elasto-capillary interactions<sup>22</sup> between the liquid and Ag, the Ag nanopillars show bending deformations, as shown in Fig. 4 (Green panel) and supporting movie S1. If two nanopillars are close enough, they collapse which results in the clustering, as shown in Fig. 4. After that, the evaporation simulation was conducted, starting with the equilibrium state obtained in the simulations of the wetting process, and evaporation evolved as shown in Fig. 4 (Blue panel) and movie S2. The procedure for evaporation has been described in previous work.<sup>23</sup> Also shown in previous work is that the evaporation could lead to further bending of nanopillars,<sup>15</sup> which may enhance the clustering of Ag nanopillars. However, we could not observe such result in this evaporation case. The reason may be that the clustering in the wetting process increases the bending stiffness of Ag nanopillars. Nevertheless, these MD simulation results demonstrate that the elasto-capillary effect is crucial for the clustering of Ag pillars observed in experiments.

In conclusion, Picric acid is an interesting molecule in many aspects affecting human sustainability and living on many levels. Its high acidity and negative charge in most natural environments presents analytical challenges, especially for plasmonic based sensing with gold or silver substrates. We demonstrate the first example of ultra-sensitive SERS of Picric acid on bare silver substrates. Particular nanofluidics and plasmonics are benefitted from, leading to enrichment of PA molecules in SERS hot-spots. Experiments and molecular dynamics simulations show that elasto-capillarity is crucial for these signal accumulating effects. With a small and simple hand-held device sensitivity and detection is in level with or even better than state of the art techniques for Picric acid today, including mass-spectrometry. Low femtogram detection is achieved from 2 µl droplets of 40 ppt Picric acid.



## ASSOCIATED CONTENT

### Supporting Information

The Supporting Information is available free of charge on the ACS Publications website. Supporting movie S1, S2 and table S1.

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### Notes

The authors declare no competing financial interests.

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